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Candle Lab Answers

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## **Heat Combustion Candle Lab**

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## **Answers**

The Heat of Combustion of candle wax is  $-21,464 \text{ kJ/mole}$  ( $H = -21,464 \text{ kJ/mole}$ ). Calculate the number of kilojoules of heat that were produced from burning your candle. Calculate the amount of time it would take to burn one mole of candle wax. (Right now, you have the moles of candle wax consumed per

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second (this is from the lab).

## **Title: Heat of Combustion of a Candle**

You can measure the amount of heat released by using a calorimeter. The calorimeter holds water and measures changes in the water's temperature. You heat the water with a candle and

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calculate the combustion heat of the candle by using the mass and the temperature change of the water. In this investigation, you will find out how much heat it takes to raise the temperature of a certain amount of water and what the combustion heat of candle wax is.

**River Dell Regional School District /**

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## **Homepage**

Heat of Combustion of Candle Wax

Purpose: To observe a burning candle and calculate the heat associated with the combustion reaction. Equipment &

Materials: • balance • candle • matches  
• modeling clay • ruler • weigh boat,

small Procedure: 1. Measure and record the length of a candle in centimeters. 2.



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## **LAB - Candle Heat of Combustion**

In this laboratory, you will burn a candle to heat up a soda can containing some water. By measuring the temperature change of the water, and the mass of candle burned, you will be able to determine the energy released when candle wax is burned on a kJ/mol basis.

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That is, you will determine the heat of combustion of candle wax.

## **Heat of Combustion of Candle Wax - Science Done Wright**

Once you calculate  $Q$ , this approximates the heat released by burning the candle.  
heat release =  $Q/dm$  units are cal/gm or kcal/gm.  $dm$  = change in mass of the

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candle. or use  $Q/\text{change in moles} = Q/0.74\dots$

## **combustion of a candle lab HELP PLEASE!? | Yahoo Answers**

the procedure of the lab is as follows: 1) determine the initial mass of candle 2) place the large can open at both ends, over the candle 3) fill 200g of water into

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a small can 4) record the...

## **Chemistry 11: combustion of a candle lab - Yahoo Answers**

Candles release heat into its surroundings. The objective of this lab was to observe and determine how much energy is released when candle wax and oxygen react to form carbon

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dioxide and water. We calculate the amount of energy transferred to the water from the candle with the amount of wax burned in our experiment. Mass of Candle Burned:

## **Candle Lab by Karoline Lin on Prezi Next**

between the initial mass of the candle

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and the final mass of the candle. 5. Using the formula below, calculate the paraffin's heat of combustion in J/gram. 6. Convert the heat of combustion of paraffin into kJ/gram. 1 kilojoule = 1000 joules. 7. The actual heat of combustion of paraffin is approximately 42 kJ/gram. Using this

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## **Heat of Combustion of Paraffin Wax**

Lab: Determining the Molar Heat of Combustion for Paraffin Procedure: 1. Note: do not light the candle until you are ins. ... Read and Download Ebook Gizmo Calorimetry Lab Answers Key PDF at Public Ebook Library GIZMO CALORIMETRY LAB ANSWERS KEY. calorimetry key . Calorimetry Wkst .

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CH301 LaBrake and Vanden Bout  
Calorimetry 1. A calorimeter was ...

## **calorimetry candle lab - PDF Free Download**

Data: Mass of Candle before burning:  
8.05g volume of water used: 200.00mL  
initial temperature of water before  
heating: 16.0 degrees Celsius room



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temperature: 24.0 degrees Celsius final  
temperature:...

## **Check over my paraffin wax lab? (heat of combustion ...**

Using an insulator that directs the heat flow to the can with the water to reduce the heat being used to burn the candle. Ensure that the cold water in the

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experiment used is near the same temperature at the starting temperature for all trials to reduce range in temperature which will affect the amount of heat absorbed.

## **Candle Lab by Alec Camp on Prezi Next**

Divide the the thermal energy created

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by mass (in grams) of the candle that were burned to calculate the heat of combustion of paraffin wax expressed in J/g.

## **What is the heat of combustion of wax - Answers**

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1. The metal can is a good thermal conductor and transfers and absorbs heat from the candle and the water and releases it to the atmosphere. This causes the calculated heat per gram to decrease. 2.

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this experiment you will determine how much energy is released when ...

## **Answers For The Heat Of Combustion Lab - Free PDF File Sharing**

Wax itself is the candle fuel. The heat from the candle melt and vaporized the wax to fuel gas and move up to

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combustion zone. Experiment to try: do it under adult supervision. Heat the wax with...

## **What is the molar heat of combustion of candle wax? - Answers**

The teacher provided the following chemical equation to describe the



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burning, or combustion, in the candle:  
Hydrocarbon fuel + O<sub>2</sub> → CO<sub>2</sub> + H<sub>2</sub>O  
The candle we used was a small, white  
emergency candle about 3.5 cm tall. The  
candle had already been used by  
someone else because I could see from  
the black wick that it had been lit before.

## **Combustion of a Candle Lab**

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## **Research Paper - 1468 Words**

The heat transferred to solution is calculated:  $q = m * c_p * T = 100 * 0.00418 \text{ kJ / g} * \text{K} * 6.4 \text{ K} = 2.68 \text{ kJ}$  view the full answer

## **Solved: Reaction 2 HEAT OF COMBUSTION: MAGNESIUM DATA AND ...**

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Moles And Chemical Formulas Lab 11  
Answers

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