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Gas (# of moles of gas is constant) Gas  
# of moles of Gas in container (mol)  
Volume of Container (L) Temperature of  
Gas (K) Pressure of Gas (kPa) A 0.3008  
14.0 672.0 120.0 B 0.3008 16.0 1280  
200.0 C 0.3008 10.0 340.0 85.0 Table 2:  
Experimental Data for Three Different  
Samples of Gases That Contain Different  
Amounts of Gas (# of moles of gas is ...

## **Information: The Gas Constant**

Chem 1A-Lec 12-Moles Part 1  
CHEMQUEST 02 [CBSE AIPMT 2007]  
Concentrated aqueous sulphuric acid is

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98%  $\text{H}_2\text{SO}_4$  by mass and has a density of  $1.80 \text{ g mL}^{-1}$ . Volume of CHEMQUEST 01 A mixture of two gases A and B in the mole ratio 2:3 is kept in a 2 litre vessel. a second 3 litre vessel has the same two gases in the

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2:3 is kept in a 2 litre vessel. a second 3 litre vessel has the same two gases, in the Intro to Chemistry 7.1: Intro to Gases

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3. On planet X, 2.78 moles of a gas takes up 1.85 L under a pressure of 74.1 kPa and a temperature of 201°C. What is the value of the ideal gas constant (R) on planet X? (include units) 0.104 kPa-L/(mol-K)

4. At a pressure of 103 kPa and a temperature of 22°C, 52.9 g of a certain gas has a volume of 31.5 L. What is the identity of this gas?

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8 Information: Structure of the Atom  
Note the following symbols: (they are  
not to scale) 35 30 50 61 Hour

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contains 3 moles of hydrogen gas and 2  
moles of oxygen gas The total pressure  
in the container is Chemquest 36

Answers LTHS Answers Skill Practice 36

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_ 1

Under water where the temperature is  
17°C and the pressure is 394 kPa, a  
diver inhales 21 L of air

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