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CHEMISTRY READING GUIDE CHAPTER
12 STOICHIOMETRY Name _____ per ____

12.1 What is Stoichiometry? 1. What is meant by the use of the term “quantitative’ in the definition of stoichiometry on page 354? 2. How does conservation of mass have to do with

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stoichiometry? 3. Explain why there is no conservation of moles in a balanced chemical ...

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Chapter 12 Stoichiometry127 SECTION
12.1 THE ARITHMETIC OF EQUATIONS
(pages 353–358) This section explains
how to calculate the amount of
reactants required or product formed in
a nonchemical process. It teaches you

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how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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Chapter 12 Stoichiometry 127 SECTION

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12.1 THE ARITHMETIC OF EQUATIONS (pages 353-358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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128 Guided Reading and Study Workbook CHAPTER 12, Stoichiometry (continued) 8 The coefficients of a balanced chemical equation tell you the

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relative number of moles of _____ and _____ in a chemical reaction 9 Why is the relative number of moles of

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